

Characterization and Atmospheric Implication of Hydrotrioxy Radical-Water-Methylamine-Formic Acid-Sulphuric Acid Complexes

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Abstract

New particle formation is an important source of atmospheric aerosols, but the nucleation phenomena is still poorly understood. Here the formation of bimolecular complexes of $\text{HOOO}^{\bullet}$ radical with H_2O , CH_3NH_2 , HCOOH and H_2SO_4 has been investigated by quantum chemical methods. The stabilising effect of methylamine is found to be close to that of sulphuric acid, both complexes are formed spontaneously at studied atmospheric conditions. In addition, the hydration of the four most stable 1:1 complexes forming the 1:1:1 complexes have been considered. Water stabilises a $\text{HOOO}^{\bullet}\cdots\text{H}_2\text{SO}_4$ complex, the most effectively. Complexation process is spontaneous as well. Complex formations induced significant red-shift and enhancement of the IR intensity for the OH stretching vibration relative to that of the free radical. Equilibrium constants have been examined and the relative abundance of complexes in the troposphere is discussed.

Keywords: hydrotrioxy radical, atmospheric acids, amine, water, nucleation precursors, quantum chemistry

1. Introduction

Hydrotrioxy ($\text{HOOO}^{\bullet}$) radical, or hydridotrioxygen according to IUPAC nomenclature¹, belongs to the group of polyoxides with general molecular formula, RO_nR , where R stands for hydrogen, any other atom or chemical group and $n \geq 3$.^{2,3} It is believed that such

polyoxides, especially the HOOO[•] radical, play an important role in atmospheric processes in Earth's troposphere and environmental reactions as well as in biological oxidation reactions.⁴ Hence the understanding of its structure, stability and reactivity with other atmospheric species is a matter of the utmost importance.

The HOOO[•] radical can be regarded as an adduct of HO[•] and O₂, and represents a temporary reservoir of OH[•] radicals.⁵ Many theoretical and experimental data suggest that the formation of HOOO[•] radical in the atmosphere can be described by the following reactions^{5,6}:



Reaction 3 is exothermic with the activation energy of around 3 kcal mol⁻¹ implying that a certain amount of the radical is indeed formed.⁶

The first mention of the hidrotrioxo radical dates back to the 1970's and ever since it has been a subject of *pomum Eridis*.⁷ Despite its long postulated existence, the HOOO[•] radical was inadvertently predicted in 1996 by Speranza⁸ and was first detected three years later by Cacace et al.⁹ in the gas phase experiments based on neutralization-reionization/collisionally activated dissociation mass spectrometry. At the turn of the millennium, Nelander et al.¹⁰, investigated and detected HOOO[•] radical in argon matrices using infrared spectroscopy. A few years later, in 2005 Suma et al.¹¹ reported the rotational spectrum and structure of the radical. Their experimental study was the first to confirm the planar geometry of the radical, suggesting *trans*-HOOO[•] conformation with inner O–O bond length of 1.688 Å.

Numerous theoretical studies on the structure and stability of the radical have been published, but there is still no consensus on its planar structure (*cis*-HOOO[•] or *trans*-HOOO[•]) as the quantum chemical calculations are strongly dependent on the level of theoretical method and basis sets used in the calculations.^{1,4,12,13,14}

HOOO[•] is a reactive molecule and can form complexes by interaction with other species present in the atmosphere. Recent works of Kim et al.¹⁵ show that the hydrated radical may be 20 to 500 times more stable compared to the isolated HOOO[•] structure. Furthermore, Cannon et al.¹⁶ have reported that the radical forms a stable binary and ternary complexes with water and sulphuric acid, which are up to 800 times more stable than the isolated radical.

Water is the most abundant species in the atmosphere, its concentration varies between 1–4 %¹⁷, and can form strong hydrogen bonds with other molecules and radical species. Such interactions can lead to the formation of new atmospheric particles that serve as a predecessors of secondary aerosols.^{18,19} In addition to the water, sulphuric acid is the main

molecular species in new particle formation.¹⁹ Many recent studies reported the importance of gaseous amines (especially methyl-, dimethyl- and trimethylamine) in nucleation processes.¹⁶ Nucleation is enhanced also by organic compounds, formic acid is an important trace constituent of the troposphere.¹⁸ The abundance of these species varies, but experimental data proposes that the concentration of sulphuric acid and methylamine is hovering around 1×10^7 molecule cm^{-3} ^{18,20}, while a concentration of formic acid is higher, around 2×10^{11} molecule cm^{-3} ¹⁸. Their concentration can significantly increase in urban and industrial areas (up to a few orders of magnitude) and varies pretty much by daytime and season.²⁰

In this paper we have performed a study on electronic structure calculations of possible conformations of the hydrotrioxy radical and its complexes with water, sulphuric acid, methylamine and formic acid molecules. Further, the hydration of the most stable 1:1 complexes were investigated. Based on these calculations, we reported valuable properties of complexes, such as the equilibrium structure, stability, the nature of interaction between each component and the IR spectral features. Thermodynamic properties of the clusters were calculated at four temperatures relevant to the atmospheric conditions. The computed Gibbs free energies were used to estimate the relative concentration of the formed complexes.

2. Computational Methods

All calculations were performed out with the GAUSSIAN 16 software package²¹ using default convergence criteria. Equilibrium geometry of monomers and complexes were optimised using the Becke three-parameter non-local exchange functional with the non-local correlation of Lee, Yang, and Parr (B3LYP)^{22,23} and the People-type 6-311++G(3df,3pd)²⁴ basis set. The harmonic and anharmonic frequencies of all species were computed at the same level of theory in order to confirm the nature of the stationary points. The final energies of the complexes were improved by a multi-step composite method using the CBS-QB3 level of theory.²⁵ The selection of above methods was based on the comparison of test calculations for hydrotrioxy radical, and available experimental structural, spectroscopic and binding parameters. These calculations included B3LYP, M06-2X²⁶ density functional methods and CCSD(T) wave function method²⁷ in combination with the 6-311++G(2df,2pd), 6-311++G(3df,3pd) and aug-cc-pVDZ basis sets. The results of these computations are summarised in the next section.

Binding energy of the complexes has been estimated as a negative difference between the zero-point corrected energy of the complexes and monomers. Specifically, for three-body complex the equation is written as:

$$D_0 = -(E_{0, \text{complex}} - E_{0, \text{monomer1}} - E_{0, \text{monomer2}} - E_{0, \text{monomer3}}). \quad (1)$$

Thermochemical parameters, i.e. Gibbs free energies, are calculated at four atmospherically relevant conditions: at the surface, 298.15 K and 278.15 K, 1atm; at altitude 5km, 255.15 K, 0.533 atm and at altitude 8 km, 235.15 K, 0.351 atm. Further, the equilibrium constants, K_c , of the complex formations were calculated according to equation 2:

$$K_c(T) = \sigma \frac{10^3 RT}{N_A} e^{-\frac{\Delta G_T}{RT}}. \quad (2)$$

In this equation σ is the symmetry number, R is the universal gas constant, T is the temperature, N_A is the Avogadro's number and ΔG_T is the complex formation Gibbs free energy.

3. Results and Discussion

Let us first consider the $\text{HOOO}^\bullet$ radical, the optimised structures of *cis* and *trans* conformations denoted as c-HTO and t-HTO, respectively, are presented in Figure 1 with the atom enumeration. Our calculations confirm that both conformations exhibit planar geometry. Due to a very low isomerization energy barrier the rotation around inner $\text{O}^2\text{--O}^3$ bond is nearly spontaneous.

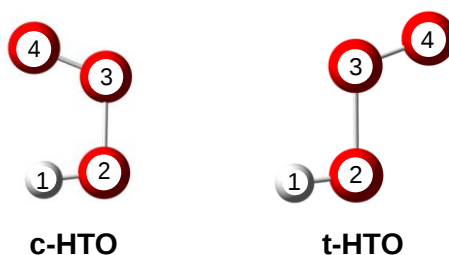


Figure 1: Optimised structures of *cis*-HOOO $^\bullet$ (c-HTO) and *trans*-HOOO $^\bullet$ (t-HTO) configuration of radical.

Structural parameters obtained at different levels of theory are summarised in Table 1. Calculated values for bond lengths and angles are in good agreement with available experimental data also presented in the Table. In general, c-HTO has longer bond lengths and

narrower angles in comparison to the same parameters in t-HTO. Interestingly, for inner O^2-O^3 bond length we detected largest differences between calculated and experimental values, absolute deviation is between 0.14 Å and 0.25 Å. Calculations at B3LYP/6-311++G(3df,3pd) show good agreement with available experimental data for t-HTO, even for the most variable inner bond length, with a discrepancy of around 8 %.

Table 1: Optimised geometry parameters for c-HTO and t-HTO. Bond distances are given in Å and angles in degrees. For comparison, experimental data are included.

Method/basis set:	c-HTO				
	$d(H^1O^2)$	$d(O^2O^3)$	$d(O^3O^4)$	$a(H^1O^2O^3)$	$a(O^2O^3O^4)$
M06-2X/aug-cc-pVTZ	0.973	1.434	1.251	100.8	111.6
B3LYP/aug-cc-pVTZ	0.974	1.505	1.250	99.9	112.7
B3LYP/6-311++G(2df,2pd)	0.973	1.503	1.249	100.0	112.7
B3LYP/6-311++G(3df,3pd)	0.973	1.501	1.247	99.9	112.7
CCSD(T)/ aug-cc-pVTZ	0.973	1.500	1.241	99.8	111.9
Method/basis set:	t-HTO				
	$d(H^1O^2)$	$d(O^2O^3)$	$d(O^3O^4)$	$a(H^1O^2O^3)$	$a(O^2O^3O^4)$
M06-2X/aug-cc-pVTZ	0.969	1.458	1.240	100.0	109.4
B3LYP/aug-cc-pVTZ	0.971	1.546	1.233	98.5	110.2
B3LYP/6-311++G(2df,2pd)	0.970	1.454	1.232	98.6	110.2
B3LYP/6-311++G(3df,3pd)	0.970	1.544	1.230	98.6	110.2
CCSD(T)/ aug-cc-pVTZ	0.970	1.542	1.231	100.2	109.1
experiment ^a	0.913	1.684	1.235	92.4	110.7
experiment ^b	0.972	1.688	1.225	90.0	111.0
experiment ^c	0.970	-	1.208	-	-

^a McCarthy et al., FTMW spectrometry¹, ^b Suma et al., FTMW spectrometry¹¹, ^c Huber et al., Photoionization mass spectrometry²⁸

Calculated binding energies for both stable conformations of the radical are represented in Table 2. The energy difference between them is rather small, but D_0 is strongly dependent on the method and basis set used for calculations. Results computed at M06-2X/aug-cc-pVTZ level show no correlation between experimental and computed value. Pople’s B3LYP method, CCSD(T) method and CBS-QB3 composite method show a stronger correlation with the experiment, with the latter being the most reliable. At the CBS-QB3 level of calculation t-HTO is slightly more stable, by 0.12 kcal mol⁻¹, than the *cis* conformation. The predicted energy difference is at the limit of computational accuracy, and for this reason, we decided to take both stable conformations into account in our study.

Table 2: Binding energies (D_0) in kcal mol⁻¹ for c-HTO and t-HTO.

Method/basis set:	c-HTO	t-HTO
M06-2X/aug-cc-pVTZ	1.23	1.36
B3LYP/aug-cc-pVTZ	3.07	3.20
B3LYP/6-311++G(2df,2pd)	2.82	2.94
B3LYP/6-311++G(3df,3pd)	2.80	2.88

CCSD(T)/ aug-cc-pVTZ	2.71	2.83
CBS-QB3	2.69	2.81
experiment ^a	-	2.94

^a Picard et al., CRESU experiment²⁹

Harmonic and anharmonic frequencies for both configurations of the radical along with the IR absorption intensities at the B3LYP/6-311++G(3df,3pd) level of theory are presented in Table 3 and graphically illustrated in Figures S1 and S2 of Supplementary Material. The radical has six different vibrational modes. The comparison between calculated harmonic and anharmonic vibrational frequencies shows that OH stretching absorption peaks for t-HTO and c-HTO differ for 160 cm⁻¹ and 166 cm⁻¹, respectively; while all the others are shifted to higher energies on average by 24 cm⁻¹ in t-HTO and 31 cm⁻¹ in c-HTO. Calculated anharmonic frequencies for t-HTO and c-HTO show reasonable agreement with observed gas-phase spectra. Predicted HO stretching peak in t-HTO appears 13 cm⁻¹ lower than experimentally observed; implying that the calculated OH stretching mode for the complexes should be adequate. The differences between anharmonic and experimental values for other vibrations are slightly more significant.

Table 3: B3LYP/6-311++G(3df,3pd) calculated harmonic and anharmonic frequencies (ν) in cm⁻¹ and IR absorption intensities (I) in km mol⁻¹ of c-HTO and t-HTO. Available experimental values are given.

	Mode:	harmonic		anharmonic		experiment ^a
		ν	I	ν	I	ν
t-HTO	H ¹ O ² srt.	3742	70	3582	59	3569
	O ³ O ⁴ str.	1389	22	1353	36	-
	H ¹ O ² O ³ bend.	1239	136	1199	113	1202
	O ² O ³ O ⁴ bend.	659	64	636	53	482
	O ² O ³ str.	460	76	445	74	244
	H ¹ O ² O ³ O ⁴ tor.	183	109	193	98	129
c-HTO	H ¹ O ² srt.	3687	49	3521	42	-
	O ³ O ⁴ str.	1413	72	1362	72	-
	H ¹ O ² O ³ bend.	1240	12	1219	8	-
	O ² O ³ O ⁴ bend.	715	150	691	138	-
	O ² O ³ str.	463	19	450	22	-
	H ¹ O ² O ³ O ⁴ tor.	230	106	163	93	148

^a Murray et al., values for isolated radical in gas-phase³⁰

3.1 Structure, intramolecular interaction and binding energy

Binary complexes

148 A systematic sampling was used to search for global minima of the complexes. Nearly a
 149 hundred initial guess structures were generated where all rational interaction of radical with
 150 water (W), methylamine (MA), formic acid (FA) and sulphuric acid (SA) molecules were
 151 considered. After full geometry optimization and on the basis of the Boltzmann probability
 152 distributions two stable low-energy structures for the HTO-W, two for HTO-MA, six for
 153 HTO-FA and three for HTO-SA complexes were identified. All these structures are shown in
 154 Supplementary Material, Figure S3, while the most stable structure for each type of complex
 155 is illustrated in Figure 2. The binding energies D_0 , are given in the figure to present the
 156 complex's stability.

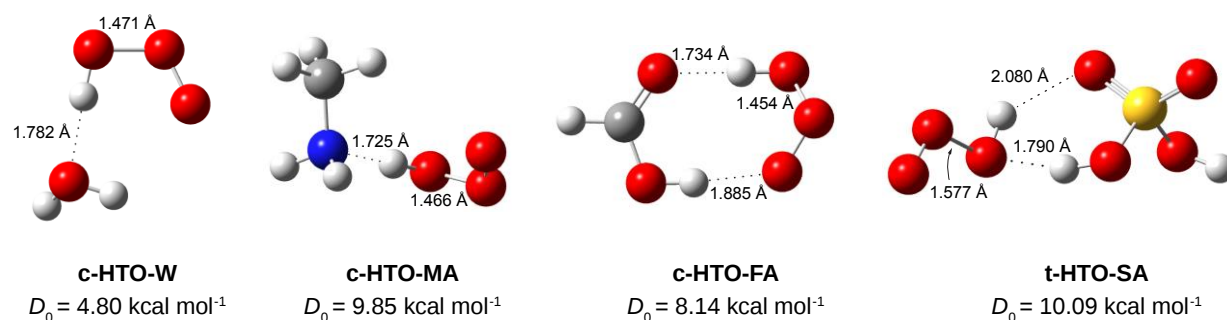


Figure 2: Optimised structures of the most stable c-HTO-W, c-HTO-MA, c-HTO-FA and t-HTO-SA complexes with the B3LYP/6-311++G(3df,3pd) bond lengths and the CBS-QB3 binding energies. Dotted lines represent intermolecularly formed H-bonds.

c-HTO-W complex. After complexation a planar structure of radical moiety in the complex is preserved. The inner O–O bond in the radical unit is shortened by 0.03 Å compared to that in the isolated radical. The hydrogen atom from the radical and the electron rich oxygen atom from the water molecule form one strong H-bond with the length of 1.782 Å. Our calculated binding energy for this structure is $4.80 \text{ kcal mol}^{-1}$, which is in good agreement with the previously reported value of $4.86 \text{ kcal mol}^{-1}$ by Kim et al.¹⁵

c-HTO-MA complex. The formation of a binary complex between c-HTO and MA forced the radical out of a planar geometry. The inner O–O bond in the c-HTO-MA complex is shortened by 0.04 Å in comparison to the isolated radical. The hydrogen bond, 1.725 Å in length, is formed between the hydrogen atom from the radical and the nitrogen atom of the methylamine. Binding energy of this complex is $9.85 \text{ kcal mol}^{-1}$.

c-HTO-FA complex. Due to the interaction between c-HTO and FA, the radical moiety undergoes a structural change, its geometry in the complex is no longer planar and the inner bond is shortened by 0.05 Å. The complex is found to be of a cyclic structure with two H-bonds. A shorter hydrogen bond, 1.734 Å, is formed between the carbonyl oxygen atom of

the formic acid and the hydrogen atom of the radical, but the radical can also act as a proton acceptor, resulting in a second H-bond, 1.885 Å, with the acid proton. The calculated binding energy for this complex is 8.14 kcal mol⁻¹.

t-HTO-SA complex. When *t*-HTO interacts with SA, the inner O–O bond in the HOOO moiety of the complex is by 0.03 Å shorter than that in the free radical. Two H-bonds with the length of 1.790 Å and 2.080 Å are formed resulting in a cyclic structure with a planar *t*-HTO unit in the complex. In the shorter H-bond SA acts as a proton donor while in the second, rather longer H-bond, the hydroxyl oxygen atom from the radical interacts with a lone oxygen electron pair of SA. Our calculated binding energy for this structure is 10.09 kcal mol⁻¹, which is in good agreement with the previously reported value of 10.00 kcal mol⁻¹ by Cannon et al.¹⁶

Ternary complexes

In addition, we carried out an extensive search of a possible structure of the ternary complexes formed by hydration of the four stable binary complexes described above. Around two hundred initial guess structures were full geometry optimised and three, five, three and six stable low-energy structures of the HTO-W, HTO-MA, HTO-FA and HTO-SA complexes with water, respectively, were selected. The identified most stable structures of the HTO-W-W, HTO-MA-W, HTO-FA-W and HTO-SA-W are shown in Figure 3, whereas all structures are found in Figure S4 of the Supplementary Material. The binding energies D_0 and $D_{0,W}$ displayed in the figure are related to the monomer interaction process and the process of hydration of corresponding binary complex, respectively.

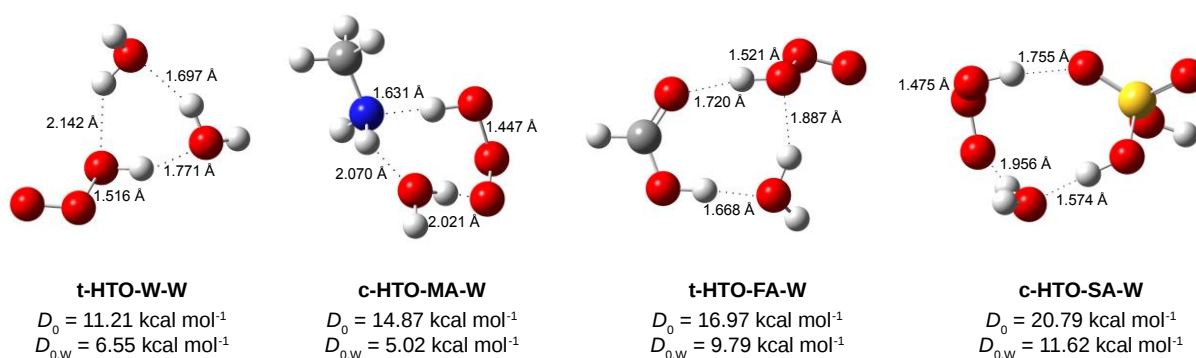


Figure 3: Optimised structures of the most stable *t*-HTO-W-W, *c*-HTO-MA-W, *t*-HTO-FA-W and *c*-HTO-SA-W complexes with the B3LYP/6-311++G(3df,3pd) bond lengths and the CBS-QB3 binding energies. Dotted lines represent intermolecularly formed H-bonds.

t-HTO-W-W complex. In the most stable ternary complex the conformation of the hydrotrioxy radical unit is changed to the *trans* arrangement compared to that in the most stable binary structure being *cis* (c-HTO-W). The inner O–O bond of the radical subunit is approximately 0.03 Å shorter than that in the isolated radical. Three H-bonds are formed by the association of the radical with two water molecules. The shortest H-bond with a length of 1.697 Å is formed between two water molecules. Next, the 1.771 Å long H-bond is a result of an interaction between the radical proton and the lone electron pair of oxygen from the water molecule. The longest H-bond at 2.142 Å results from an interaction of the hydroxyl radical oxygen atom being an electron acceptor and water hydrogen being a proton donor. t-HTO-W-W is quite stable with a binding energy of 11.21 kcal mol⁻¹. When comparing energies of the ternary complex and the corresponding binary one, we observe that additional an water molecule contributes around 6.55 kcal mol⁻¹ to the stability of the t-HTO-W-W complex.

c-HTO-MA-W complex. *Cis* configuration of the radical moiety is retained in the hydrogenation of the HTO-MA complex. The radical inner O–O bond is again shortened by 0.05 Å. The c-HTO-MA-W complex is a cyclic structure with three H-bonds. The shortest intermolecular H-bond at 1.631 Å results from an interaction between the radical hydrogen and the electron rich amino nitrogen atom. The water molecule acts as an H-donor to the terminal radical oxygen atom, and as an H-acceptor from the amine group, the formed H-bonds are 2.021 and 2.070 Å long, respectively. Further, calculated binding energy for this complex related to the monomer interaction process is 14.87 kcal mol⁻¹, whereas the additional water contributes 5.02 kcal mol⁻¹ to the stability of the complex.

t-HTO-FA-W complex. The hydrotrioxy radical moiety in the complex changes the orientation of atoms when the water molecule is added to the binary HTO-FA complex. The inner O–O bond is shortened by 0.02 Å. Three stable intermolecular bonds are formed, resulting in an eight-membered ring structure. The water molecule acts as a proton donor and proton acceptor and forms two 1.887 Å and 1.668 Å long H-bonds, respectively. The lone electron pair on the carbonyl oxygen atom participates in the 1.720 Å long H-bond with the hydrogen atom from the HOOO subunit. The binding energy is predicted to be 16.96 kcal mol⁻¹. Additional water stabilises the most stable HTO-FA complex even further, by 9.06 kcal mol⁻¹.

c-HTO-SA-W complex. Similar to the t-HTO-W-W and t-HTO-FA-W complexes, the radical moiety changes the orientation in comparison to the corresponding binary complex. As in all the other cases, the inner O–O bond in the HTO unit is shortened, in this case by 0.03 Å. A strong H-bond, 1.574 Å long, is formed between the hydrogen atom of SA and the lone

electron pair on the water oxygen atom, whereas the H-bond with the acid being a proton acceptor is 1.755 Å long. The third intermolecular H-bond is longer, 1.956 Å, where water acts as the proton donor and terminal oxygen contributes a lone electron pair to the bond. The c-HTO-SA-W complex with the binding energy of 20.79 kcal mol⁻¹ is predicted to be the most stable ternary structure reported in this study. Water stabilises the HTO-SA significantly, as it contributes 11.62 kcal mol⁻¹ to the stability of the c-HTO-SA-W complex.

3.2 Vibrational spectra

The complete set of harmonic and anharmonic frequencies with IR absorption intensities calculated at B3LYP/6-311++G(3df,3pd) level of theory for the both HOOO[•] conformers and the precursor molecules is collected in Table S1 of the Supplementary Material, where available experimental data for gas-phase molecules of precursors are given for comparison. Figures 4 and 5 show harmonic vibrational spectra for the four most stable binary and the four most stable ternary complexes, respectively.

The intermolecular interactions in complex formation induced new degrees of freedom and consequently new intermolecular vibrational modes which are not present in monomer spectra. An examination of the calculated data shows that the frequencies and IR absorption intensities of H-bonded OH stretching regions are the most significantly affected by complexation. As expected, a large frequency red-shift is mainly observed for vibration of OH bonds involved in H-bond formation.

Binary complexes

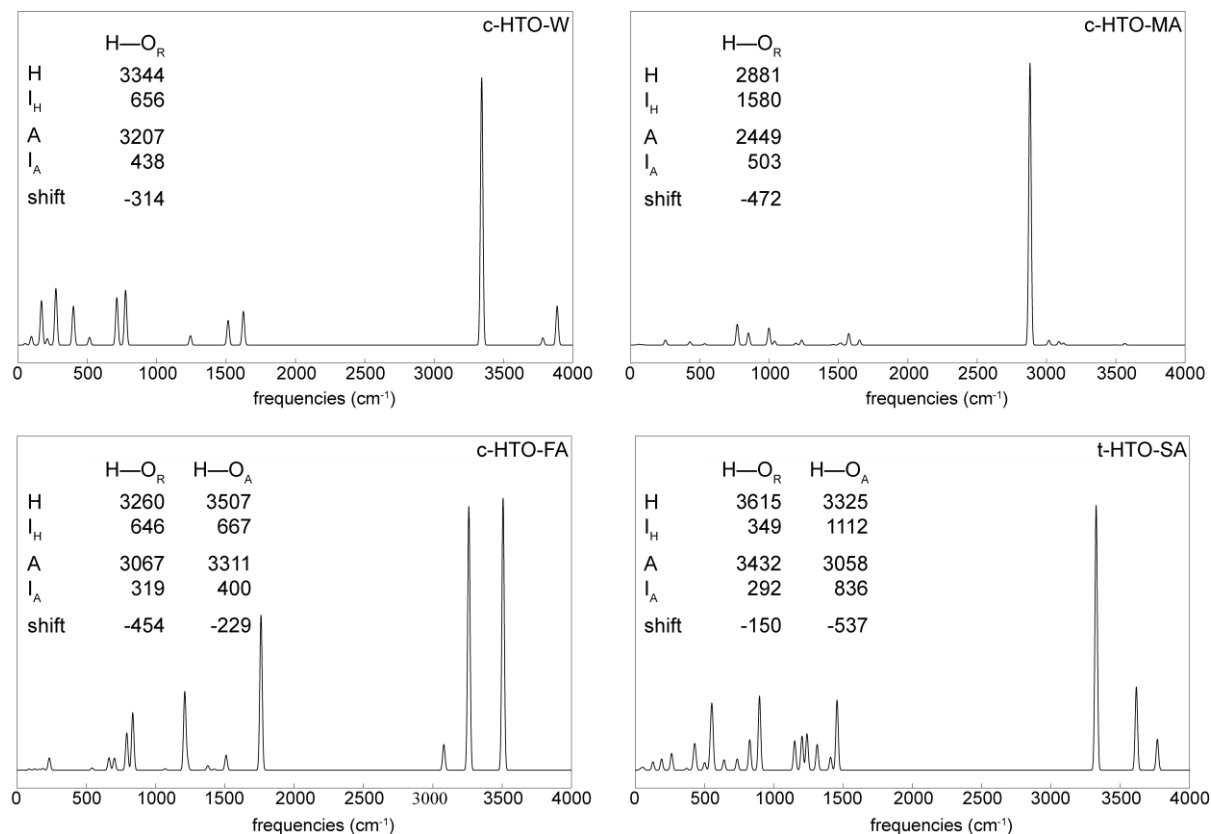


Figure 4: B3LYP/6-3++1G(3df,3pd) calculated IR absorption spectra for the most stable binary complexes. Graphs represent the harmonic absorption spectra, while in the inserted tables the relevant harmonic and anharmonic absorption peaks as well as calculated shifts of anharmonic frequencies in cm^{-1} are summarised.

c-HTO-W complex. The complex has 15 fundamental vibrational modes, 6 of them are related to the free radical and three are related to modes in an isolated water molecule. The remaining six are present due to intermolecular interactions. The most noticeable change in the IR spectrum is a red-shift of 314 cm^{-1} related to the OH stretching of the $\text{HOOO}^{\bullet}$ radical subunit. This stretching vibration is decreased to 3207 cm^{-1} in the complex relative to the 3521 cm^{-1} in the isolated c-HTO, while the intensity of the transition increases to the maximum in the spectra of the complex.

c-HTO-MA complex. This particular complex has even more vibrational modes due to more participating atoms and bonds in the complex formation. When the OH group from the radical subunit in the complex is compared to the isolated $\text{HOOO}^{\bullet}$, it is seen that frequency is red-shifted by 472 cm^{-1} and is accompanied by a substantial increase in IR absorption.

c-HTO-FA complex. When the anharmonic vibrational frequencies of the OH symmetric stretch mode of cyclic c-HTO-FA is considered, a shift of 454 cm^{-1} to the red region of the spectra with increased intensity is observed. The OH vibration of the formic acid subunit is also shifted to the higher energy by 229 cm^{-1} .

t-HTO-SA complex. In this complex, the OH stretching vibration of the enol HTO and sulphuric acid subunits are also red-shifted compared to the isolated monomers, the shifts are 150 cm^{-1} and 537 cm^{-1} , respectively.

Ternary complexes

In general, ternary complexes have even more vibrational modes due to the three body molecules interactions forming stable complexes.

t-HTO-W-W complex. The radical and both water OH stretching frequencies are moderately red-shifted; the shifts are 419 cm^{-1} , 198 cm^{-1} and 87 cm^{-1} compared to the OH symmetric stretching motions of the radical and water subunits. The vibrational modes are all accompanied by increased intensity.

c-HTO-MA-W complex. The most pronounced change in the vibrational spectrum of the complex with regard to the isolated species is the 419 cm^{-1} red-shift of the OH anharmonic stretching frequency in the HOOO moiety. The OH and NH stretching of the water and amine moieties are decreased only by 87 cm^{-1} and 48 cm^{-1} , respectively, implying the less strong participation of water and amino groups in the bonding.

t-HTO-FA-W complex. When the OH-stretching anharmonic frequencies of the OH-stretching in the t-HTO-FA-W complex are considered and compared to the isolated species, the red-shifts of 671 cm^{-1} , 511 cm^{-1} and 237 cm^{-1} are observed for the OH modes of the radical, formic acid and water moiety, respectively, all involved in the H-bond formation.

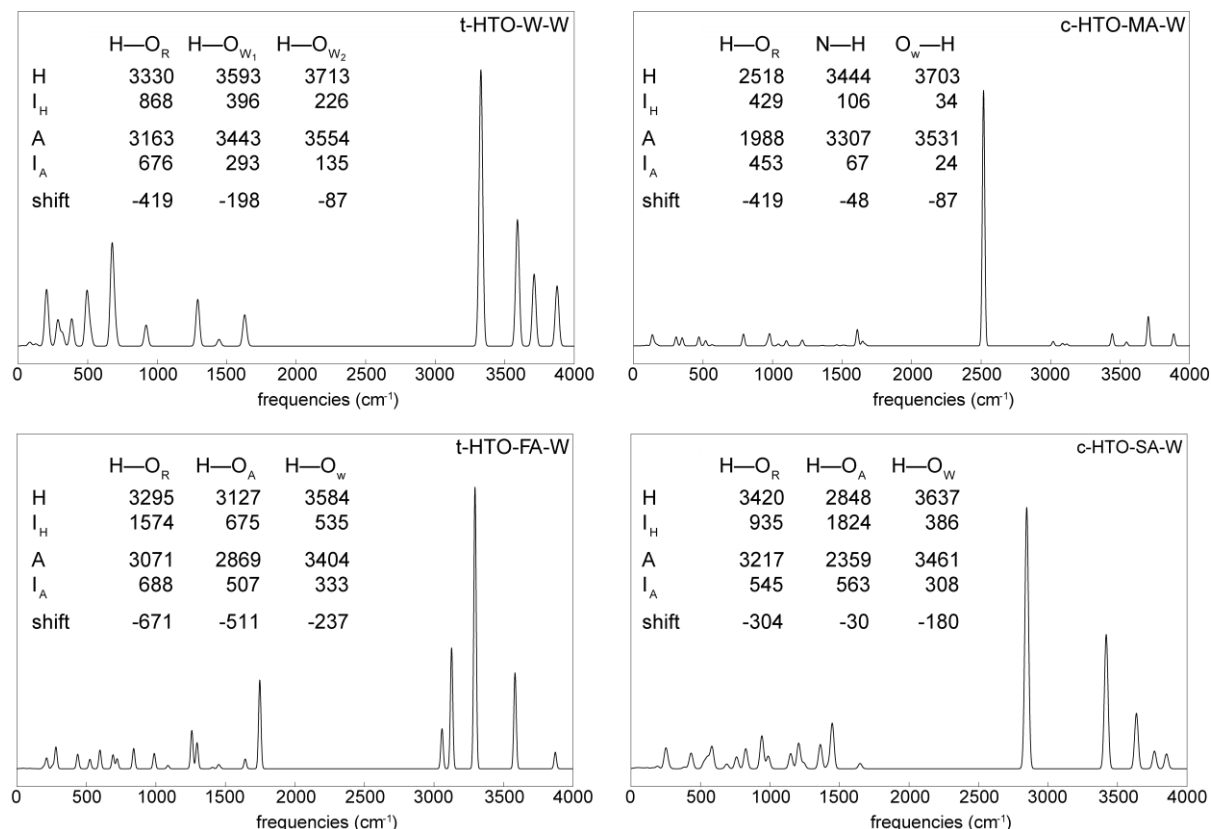


Figure 5: B3LYP/6-311++G(3df,3pd) calculated IR absorption spectra for the most stable ternary complexes. Graphs represent harmonic absorption spectra, while in the inseted tables the relevant harmonic and anharmonic absorption peaks as well as calculated shifts for anharmonic frequencies in cm^{-1} are summarised.

c-HTO-SA-W complex. In this specific complex, shifts in three OH vibrational modes are observed due to participation in the H-bond formation. The OH stretching in HOOO moiety is shifted by 304 cm^{-1} . Sulphuric acid OH bond related vibration is decreased by 30 cm^{-1} , and symmetric OH vibrational mode of water moiety is red-shifted by 180 cm^{-1} .

Evaluated data provides strong evidence that the complex formation of the hydrotrioxo radical with water, methylamine, formic acid and sulphuric acid introduces new intermolecular vibrations. The vibration frequencies of those OH bonds that participate in H-bonding interaction are generally shifted to lower frequencies, while their intensity absorption bands strongly increase. The other modes in the complexes are similar to that in the isolated monomers.

3.3 Atmospheric Implication

Binding energies are important parameters related to the stability of the complexes formed, and the vibrational spectra provide valuable data for laboratory and atmospheric identification. To provide insightful information into the atmospheric relevance of the complexes the thermochemical properties have also to be evaluated.

Table 4: Thermodynamic properties (ΔH , $T\Delta S$ and ΔG) in kcal mol⁻¹ and calculated K_c values in cm³ molecules⁻¹ at different altitude (h) [km], temperature (T) [K] and pressure (p) [atm].

Complex	h	T	p	ΔH	$T\Delta S$	ΔG	K_c
Binary Complexes							
c-HTO-W	0	298.15	1	-5.42	-8.40	2.98	2.8×10^{-22}
	0	278.15	1	-5.44	-7.84	2.40	5.1×10^{-22}
	5	255.15	0.533	-5.46	-7.53	2.07	6.1×10^{-22}
	8	236.15	0.351	-5.47	-7.18	1.71	8.8×10^{-22}
c-HTO-MA	0	298.15	1	-9.94	-9.24	-0.70	1.4×10^{-19}
	0	278.15	1	-9.97	-8.64	-1.33	4.4×10^{-19}
	5	255.15	0.533	-9.99	-8.25	-1.74	1.1×10^{-18}
	8	236.15	0.351	-10.01	-7.84	-2.16	3.4×10^{-18}
c-HTO-FA	0	298.15	1	-8.37	-10.29	1.92	1.7×10^{-21}
	0	278.15	1	-8.40	-9.63	1.23	4.3×10^{-21}
	5	255.15	0.533	-8.44	-9.19	0.75	8.3×10^{-21}
	8	236.15	0.351	-8.47	-8.72	0.26	1.9×10^{-20}
t-HTO-SA	0	298.15	1	-10.02	-9.45	-0.56	1.1×10^{-19}
	0	278.15	1	-10.06	-8.86	-1.20	3.5×10^{-19}
	5	255.15	0.533	-10.10	-8.49	-1.61	8.8×10^{-19}
	8	236.15	0.351	-10.13	-8.08	-2.05	2.7×10^{-18}
Ternary Complexes							
t-HTO-W-W	0	298.15	1	-12.66	-18.04	5.38	4.8×10^{-24}
	0	278.15	1	-12.70	-16.86	4.16	2.1×10^{-23}
	5	255.15	0.533	-12.72	-16.13	3.41	4.4×10^{-23}
	8	236.15	0.351	-12.74	-15.33	2.59	1.3×10^{-22}
c-HTO-MA-W	0	298.15	1	-15.59	-18.29	2.70	4.5×10^{-22}
	0	278.15	1	-15.65	-17.12	1.47	2.8×10^{-21}
	5	255.15	0.533	-15.71	-16.40	0.69	9.3×10^{-21}
	8	236.15	0.351	-15.75	-15.61	-0.14	4.5×10^{-20}
t-HTO-FA-W	0	298.15	1	-17.94	-19.69	1.75	2.2×10^{-21}
	0	278.15	1	-17.99	-18.42	0.43	1.8×10^{-20}
	5	255.15	0.533	-18.04	-17.58	-0.46	8.9×10^{-20}
	8	236.15	0.351	-18.07	-19.69	-1.38	6.3×10^{-19}
c-HTO-SA-W	0	298.15	1	-21.45	-18.89	-2.56	3.2×10^{-18}
	0	278.15	1	-21.50	-17.67	-3.83	4.1×10^{-17}
	5	255.15	0.533	-21.55	-16.69	-4.66	3.6×10^{-16}
	8	236.15	0.351	-21.58	-16.06	-5.53	4.4×10^{-15}

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172 The enthalpy (ΔH), entropy ($T\Delta S$) and Gibbs free energy (ΔG) changes for global minimum
 173 structures of binary and ternary complexes were calculated as a difference between the values
 174 for the complex and infinity separated constituents. These quantities at different atmospheric
 175 conditions as noted in the introduction section together with calculated K_c values according to
 176 the equation 2 are gathered in Table 4.

An inspection of the table illustrates that both the binary and ternary complex formations are exothermic processes at all studied conditions. The ΔH value decreases with increasing atmospheric height, but according to the ΔG value, not all processes are spontaneous. The c-HTO-MA and t-HTO-SA binary complexes, are spontaneously formed at all conditions with $\Delta G_{236.15}$ -2.16 kcal mol⁻¹ and -2.05 kcal mol⁻¹, respectively. Only the c-HTO-SA-W ternary complex is formed spontaneously at all studied altitudes with $\Delta G_{236.15} = -5.53$ kcal mol⁻¹, while c-HTO-MA-W and t-HTO-FA-W have a negative Gibbs free energy value at higher levels of the troposphere. K_c values of the complexes increase with height, probably due to lower pressure and lower concentration of the reactive species. From the equilibrium constants and the gas-phase concentration of molecular species, we have evaluated relative abundance of the particular binary complexes. The reported atmospheric concentrations at 298 K are as follows: $[H_2O] = 8 \times 10^{17}$ molecules cm⁻³¹⁷, $[CH_3NH_2] = 1 \times 10^7$ molecules cm⁻³²⁰, $[HCOOH] = 2 \times 10^{11}$ molecules cm⁻³¹⁸ and $[H_2SO_4] = 1 \times 10^7$ molecules cm⁻³¹⁸. For example, 0.002 % of the HOOO[•] radical is bounded into the c-HTO-W complex at the Earth's surface, while only 4×10^{-4} % of the radical form a binary complex with water at an altitude of 8 km and 100 % relative humidity. For other complexes the values of relative abundance can go as far as ten decades lower than those calculated for water.

Figure 6 illustrates the computed equilibrium constants at the different temperatures at the ground level and different altitudes of the Earth's atmosphere for the all spontaneous processes. The equilibrium constants for the formation of the bimolecular t-HTO-SA and c-HTO-MA complexes are comparable. A formation of the ternary c-HTO-SA-W complex from monomes appears to be more favourable than the hydration of binary c-HTO-SA complex. The equilibrium constant for the hydration of binary complexes is one decade lower than for the monomer interaction process resulting in the same ternary c-HTO-SA-W complex formation. Thermodynamic quantities for the hydration process of binary complexes are listed in the Table S2 of the Supplementary Material.

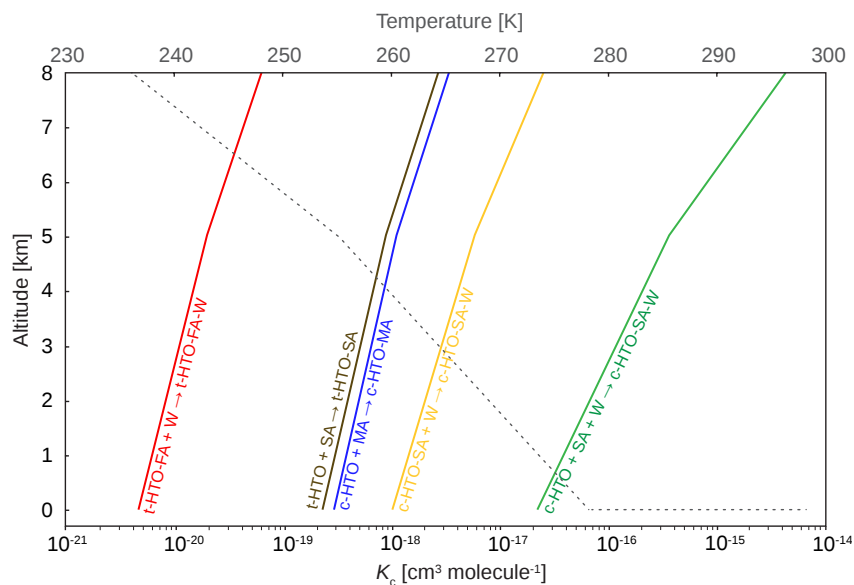


Figure 6: Calculated equilibrium constants, K_c , as a function of the altitude of Earth's atmosphere for the formation of selected processes indicated along each curve on the figure above. The dashed line represents a variation of temperature.

4. Conclusion

The hydrogen-bonded complexes of hydrotrioxy radical with water, methylamine, formic acid and sulphuric acid were studied using the B3LYP/6-311++G(3df,3pd) and CBS-QB3 methods. Our results indicate that due to the negligible energy difference, 0.1 kcal mol⁻¹, either the *cis* or *trans* isomer is involved in the complexation. Water and methylamine formed one strong H-bond with the hydrotrioxy radical with lengths of around 1.75 Å, while the complexes with both acids have a cyclic structure with two H-bonds, the second H-bonds are somewhat longer. The binding energies of complexes with methylamine and sulphuric acid are comparable, they approach to 10 kcal mol⁻¹. The addition of water to the most stable dimer complexes resulted in multiple H-bonded cyclic structures with H-bonds of lengths between 1.6 Å and 2.1 Å and stabilised the most effectively the complex with sulphuric acid, with binding energy of 20.8 kcal mol⁻¹.

The present calculations confirmed that large spectral red-shift and enhancement of IR intensities could be used for spectroscopic identification in the atmosphere. In particular, the most substantial red-shift of 671 cm⁻¹ is observed for the H-bonded OH stretching vibrational mode of HOOO moiety in the hydrated complex involved formic acid.

From the calculated thermodynamic quantities, it should be concluded that a binary complex with methylamine and sulphuric acid, and a ternary complex with sulphuric acid are all

formed spontaneously at the Earth's surface to the height of 8 km in the troposphere. Equilibrium constants of the complexes increase with increasing altitude.

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